**Weak Acid/Bases**

1. Find pH and percent dissociation for:
   1. .2M HC2H3O2
   2. .2M HIO
   3. .2M NH3
   4. .02M HClO4
2. Find Ka if H2Se has a pH of 4.1 when it is .3M
3. Find pH of HX if a .5M solution is .1% dissociated
4. Acid HQ has a pH of 3.7 and is .05% dissociated. Find:
   1. [HQ]
   2. Ka
5. N2H4 has a pH of 11.1 Find [N2H4]
6. HBrO has a pH of 4.4 Find [HBrO]
7. NH2NH3 has a pH of 11.7 and is .2% dissociated, find the Kb.
8. Find pH when 2g HCN is dissolved in 200mL
9. Find pH when 11g KOH is dissolved in 200mL
10. Find pH when 11g NH3 is dissolved in 200 mL
11. Tell whether the chemical behaves as an acid, base, or is neutral when dissolved in water.
    1. CaCl2
    2. Li3PO4
    3. K2S
    4. LiClO4
    5. Ni(NO3)3
    6. LiI
    7. KNO2
    8. NH4Br