**Strong Acid Base Problems**

1. Find the [OH-] in .15M
   1. Calcium hydroxide
   2. Potassium hydroxide
   3. Hydrochloric acid
2. If the [OH-]=.08M Find the molarity of
   1. Potassium hydroxide
   2. Barium hydroxide
3. A student has 1.5g nitric acid dissolved in 2,500 mL. Find the pH.
4. A student dissolves 1.5g sodium hydroxide in 25 L. Find the pH
5. How many grams of perchloric acid is dissolved in 20 liters of pH 5.5 solution?
6. How many grams sodium hydroxide is dissolved in 10 L of pH 10 solution?
7. How many g barium hydroxide is dissolved in 10 L of pH 12 solution?
8. Find the pH when 20mL pH 2 solution is added to 180 mL water.
9. Find the pH when 50mL pH 2 solution is diluted to 100mL.
10. 40mL pH 12 solution is diluted to 4000mL. Find the pH.
11. How much water must be added to 90 mL pH 2.5 solution to make the pH 3.5?
12. 60mL .15M nitric acid is added to 152 mL of .06M sodium hydroxide. Find the pH.
13. If 40 mL .002 M sulfuric acid is added to 80 mL .019M KOH, Find the pH.
14. 40 mL .08M HCl is added to 160 mL .04M HNO3-. Find the pH.
15. 1.2 g NaOH is added to 3L of pH 11.0 KOH. Find the pH.