Strong Acid Base Intro

1. Write the dissociation equation for
    a. Nitric acid
    b. Sulfuric acid
    c. Barium hydroxide
    d. Sodium hydroxide
    e. Perchloric acid

2. Fill in the table

pH	pOH	$[OH^{-1}]$	$[H_3O^+]$
2.1			
	3.5		
		0.007	
			0.0015

3. Find the $[H^+]$ in
    a. 0.3 M KOH
    b. 0.3 M HCl

4. Find the $[OH^-]$ in
    a. 0.003 M $Ca(OH)_2$
    b. 0.003 M HNO_3

5. If the pH is 5.3, find $[H_3O^+]$.

6. If [HI] = 0.097 M, find $[OH^-]$.

7. If $[Ba(OH)_2]$ = 0.007 M, find $[H^+]$.