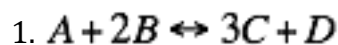
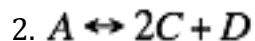


Equilibrium (1)

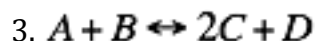


In a 3L container...

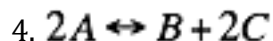
- There are 2 mol A, 1 mol B, 1.5 mol C, and 0.4 mol D at equilibrium. Find Kc
- Using Kc from (a), find [A] if [B]= 0.3 M, [C]= 0.8 M, and [D]= 0.9 M.



- Kc=0.042 Find [D] if [A] and [C] are .8M at equilibrium.
- Kc=100 Find [D] if [A] and [C] are .8M at equilibrium.

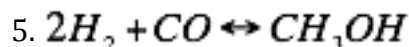


Initially in a 2L container, 0.2 mol of A and 0.3 mol of B are present. At equilibrium, 0.05 moles of A remain. Find all equilibrium values and c

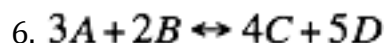


Initially in a 1L container, there is 0.8 mol of A, 0.6 mol of B, and 0.3 mol of C present.

- At equilibrium, there is 0.93 mol of B present. Find all other equilibrium values.
- Instead, at equilibrium there is 0.98 mol of A present. Find other equilibrium values.



Initially 0.6 mol H₂, 0.8 mol CO, and 0.98 mol CH₃OH were in a 2L container. At equilibrium, [CH₃OH] was 0.21 M. Find equilibrium values of all.



Initially C and D were each 0.9 M. At equilibrium, D was 0.4 M. Find Kc.